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By John Muter

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Excerpt: . There are almost unlimited cases in which such solutions may be applied to the various bodies capable of undergoing a definite change by oxidation, but the most common applications are: --. (1) For the estimation of ferrous salts. If we look at the simplest equation for the transference of iron from the ferrous to the ferric state, we find $2\text{FeO} + \text{O} = \text{Fe}_2\text{O}_3$. So we observe that 1 atomic weight of O can oxidize 2 atomic weights of Fe, or that $7.94 = \text{Fe } 55.5; 1, w .$ molecular weight therefore each Cc. of fir oxidant = 005 s 5 metallic Iron or lu JJJ 10,000 of any ferrous salt containing 1 atom of iron in its molecule. Either of the oxidants may be employed for the estimation of ferrous salts. Take, for example, crystallized ferrous sulphate. We have the equations: --With bichromate. $\text{K}_2\text{Cr}_2\text{O}_7 + 6\text{FeSO}_4 \dots$



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